

PART 1 - CHEMISTRY

Topic : Chemical kinetics

SECTION-B : (Maximum Marks : 4)

- This section contains **ONE** question.
- The answer to each question is a **NUMERICAL VALUE**.
- For each question, enter the correct numerical value of the answer (in decimal notation, truncated / rounded-off to the **SECOND DECIMAL PLACE**; e.g. 6.25, 7.00, -0.33, -.30, 30.27, -127.30) in the place designated to enter the answer.
- Answer to each question will be evaluated according to the following marking scheme :
 Full Marks : +4 If ONLY the correct answer is entered;
 Zero Marks : 0 In all other cases.

1. Consider a reaction $A + R \longrightarrow \text{Product}$. The rate of this reaction is measured to be $k[A][R]$. At the start of the reaction, the concentration of R, $[R]_0$, is 10-times the concentration of A, $[A]_0$. The reaction can be considered to be a pseudo first order reaction with assumption that $k[R] = k'$ is constant. Due to this assumption, the relative error (in %) in the rate when this reaction is 40 % complete, is _____.
 [k and k' represent corresponding rate constants]

[JEE Advanced 2025 Paper 2]

Scan or tap for the
full solution

Space for Rough Work

JEE ADVANCED CHEMICAL KINETICS — PSEUDO FIRST ORDER REACTION — PYQS

Test Pattern : JEE (Advanced)

ANSWER KEY**SECTION-B**

Ques.	1														
Ans.	4 to 4.25														