

## PART 1 - CHEMISTRY

Topic : Solid state

## SECTION-B : (Maximum Marks : 4)

- This section contains **ONE** question.
- The answer to each question is a **NUMERICAL VALUE**.
- For each question, enter the correct numerical value of the answer (in decimal notation, truncated / rounded-off to the **SECOND DECIMAL PLACE**; e.g. 6.25, 7.00, -0.33, -0.30, 30.27, -127.30) in the place designated to enter the answer.
- Answer to each question will be evaluated according to the following marking scheme :  
Full Marks : +4 If ONLY the correct answer is entered;  
Zero Marks : 0 In all other cases.

1. The density (in  $\text{g cm}^{-3}$ ) of the metal which forms a cubic close packed (ccp) lattice with an axial distance (edge length) equal to 400 pm is \_\_\_\_\_.

Use: Atomic mass of metal = 105.6 amu and Avogadro's constant =  $6 \times 10^{23} \text{ mol}^{-1}$

[JEE Advanced 2025 Paper 2]

Scan or tap for the  
full solution



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Space for Rough Work

**JEE ADVANCED SOLID STATE — DENSITY OF CUBIC UNIT CELL — PYQS**

Test Pattern : JEE (Advanced)

**ANSWER KEY****SECTION-B**

|       |               |  |  |  |  |  |  |  |  |  |  |
|-------|---------------|--|--|--|--|--|--|--|--|--|--|
| Ques. | 1             |  |  |  |  |  |  |  |  |  |  |
| Ans.  | 10.85 to 11.1 |  |  |  |  |  |  |  |  |  |  |