

JEE MAIN CHEMICAL BONDING AND MOLECULAR STRUCTURE —
DIPOLE MOMENT, POLARITY AND NON-POLAR MOLECULES — PYQS

Max. Marks: 8

Questions: 2

Section A (MCQ) — 2 questions

- Q.1** Given below are statements about some molecules/ions.
- A. The dipole moment value of NF_3 is higher than that of NH_3 .
 - B. The dipole moment value of BeH_2 is zero.
 - C. The bond order of O_2^{2-} and F_2 is same.
 - D. The formal charge on the central oxygen atom of ozone is -1 .
 - E. In NO_2 , all the three atoms satisfy the octet rule, hence it is very stable.
- Identify the CORRECT statements.
Choose the correct answer from the options given below:

[JEE Main 2026 (Online) 24th January Morning Shift]

- (1) B & C Only (2) B, C & D Only
(3) A, C & D Only (4) A, B, C, D & E

Scan or tap for the
full solution

- Q.2** Among H_2S , H_2O , NF_3 , NH_3 and CHCl_3 , identify the molecule (X) with lowest dipole moment value. The number of lone pairs of electrons present on the central atom of the molecule (X) is :

[JEE Main 2026 (Online) 22nd January Evening Shift]

- (1) 3 (2) 2 (3) 0 (4) 1

Scan or tap for the
full solution

ANSWER KEY

Ques.	1	2									
Ans.	1	4									