

JEE MAIN ELECTROCHEMISTRY — ELECTROCHEMICAL CELLS, CONSTRUCTION OF GALVANIC CELL AND CALCULATION OF STANDARD EMF — PYQS

Max. Marks: 8

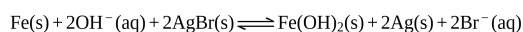
Questions: 2

Section A (MCQ) — 2 questions

- Q.1** An electrochemical cell is constructed using half cells in the direction of spontaneous change:
 $\text{Fe}(\text{OH})_2(\text{s}) + 2\text{e}^- \longrightarrow \text{Fe}(\text{s}) + 2\text{OH}^-(\text{aq})$
 $E^\theta = -0.88 \text{ V}$
 and $\text{AgBr}(\text{s}) + \text{e}^- \longrightarrow \text{Ag}(\text{s}) + \text{Br}^-(\text{aq})$
 $E^\theta = +0.07 \text{ V}$
 Which of the following options is correct?

[JEE Main 2026 (Online) 2nd April Morning Shift]

(1) Overall reaction

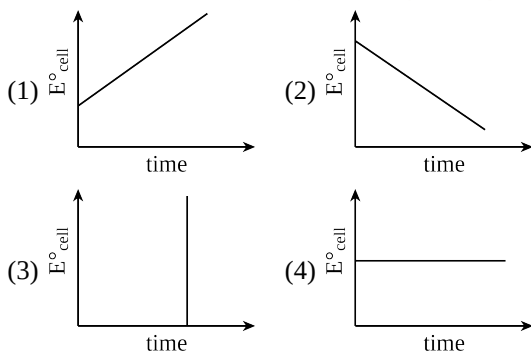
(2) $E^\theta_{\text{cell}} = -0.95 \text{ V}$

(3) Fe is reduced in the electrochemical cell

(4) E^θ_{cell} is an extensive propertyScan or tap for the
full solution

- Q.2** For a closed circuit Daniell cell, which of the following plots is the accurate one at a given temperature?

[JEE Main 2026 (Online) 21st January Evening Shift]

Scan or tap for the
full solution

ANSWER KEY

Ques.	1	2									
Ans.	1	4									