

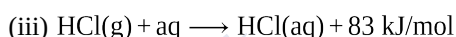
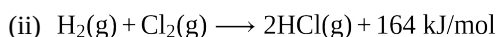
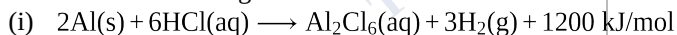
JEE MAIN THERMODYNAMICS — HESS'S LAW OF CONSTANT HEAT SUMMATION — PYQS

Max. Marks: 4

Questions: 1

Section A (MCQ) — 1 question

Q.1 Consider the following data.



The enthalpy of formation of anhydrous solid Al_2Cl_6 is :

[JEE Main 2026 (Online) 2nd April Morning Shift]

(1) -648 kJ mol^{-1} (2) $-1350 \text{ kJ mol}^{-1}$

(3) $-2002 \text{ kJ mol}^{-1}$ (4) $-1527 \text{ kJ mol}^{-1}$

Scan or tap for the
full solution

ANSWER KEY

Ques.	1									
Ans.	4									