

JEE MAIN 2026 (ONLINE) 22ND JANUARY EVENING SHIFT — QUESTION PAPER

Max. Marks: 300
Questions: 75

Section A (MCQ) — 60 questions

- Q.1** In an open organ pipe v_3 and v_6 are 3rd and 6th harmonic frequencies, respectively. If $v_6 - v_3 = 2200$ Hz then length of the pipe is _____ mm.

(Take velocity of sound in air is 330 m/s.)

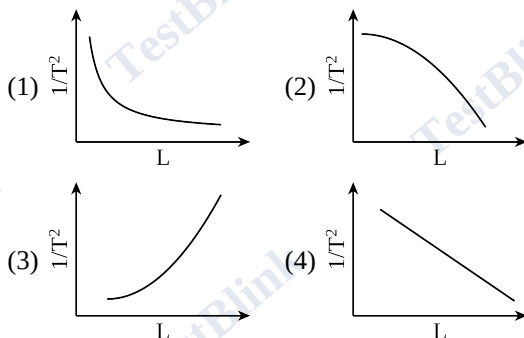
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- (1) 200 (2) 225 (3) 275 (4) 250

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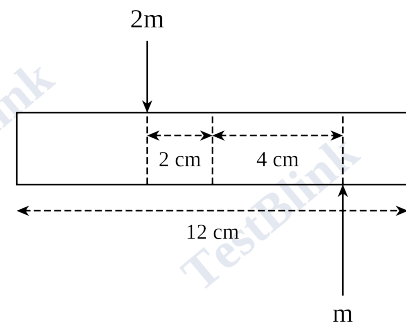
- Q.2** Using a simple pendulum experiment g is determined by measuring its time period T . Which of the following plots represent the correct relation between the pendulum length L and time period T ?

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- Q.3** A uniform bar of length 12 cm and mass 20 m lies on a smooth horizontal table. Two point masses m and $2m$ are moving in opposite directions with same speed of v and in the same plane as the bar, as shown in figure. These masses strike the bar simultaneously and get stuck to it. After collision the entire system is rotating with angular frequency ω . The ratio of v and ω is :

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- (1) 33 (2)
- $2\sqrt{88}$
- (3) 32 (4) 66

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- Q.4** A laser beam has intensity of 4.0×10^{14} W/m². The amplitude of magnetic field associated with beam is _____ T.

 (Take $\epsilon_0 = 8.85 \times 10^{-12}$ C²/Nm² and $c = 3 \times 10^8$ m/s)

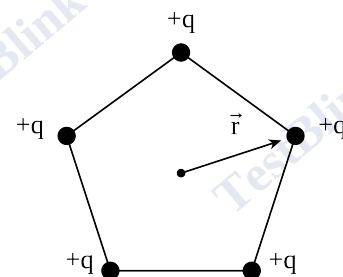
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- (1) 1.83 (2) 2.0 (3) 5.5 (4) 18.3

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- Q.5** Five positive charges each having charge q are placed at the vertices of a pentagon as shown in the figure. The electric potential (V) and the electric field ($\vec{E}$) at the center O of the pentagon due to these five positive charges are :

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$$(1) V = \frac{5q}{4\pi\epsilon_0 r} \text{ and } \vec{E} = \frac{5\sqrt{3}q}{8\pi\epsilon_0 r^2} \hat{r}$$

$$(2) V = \frac{5q}{4\pi\epsilon_0 r} \text{ and } \vec{E} = 0$$

$$(3) V = 0 \text{ and } \vec{E} = 0$$

$$(4) V = \frac{5q}{4\pi\epsilon_0 r} \text{ and } \vec{E} = \frac{5q}{4\pi\epsilon_0 r^2} \hat{r}$$

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Q.6 Given below are two statements :

Statement-I: An object moves from position r_1 to position r_2 under a conservative force field $\vec{F}$. The work done by the force is $W = -\int_{r_1}^{r_2} \vec{F} \cdot d\vec{r}$.

Statement-II: Any object moving from one location to another location can follow infinite number of paths. Therefore, the amount of work done by the object changes with the path it follows for a conservative force.

In the light of the above statements, choose the correct answer from the options given below :

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- (1) Statement-I is true but Statement-II is false
- (2) Statement-I is false but Statement-II is true
- (3) Both Statement-I and Statement-II are false
- (4) Both Statement-I and Statement-II are true

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Q.7 When a part of a straight capillary tube is placed vertically in a liquid, the liquid rises up to certain height h . If the inner radius of the capillary tube, density of the liquid and surface tension of the liquid decrease by 1% each, then the height of the liquid in the tube will change by _____ %.

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- (1) +3 (2) +1 (3) -1 (4) -3

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Q.8 Given below are two statements :

Statement-I: A satellite is moving around earth in the orbit very close to the earth surface. The time period of revolution of satellite depends upon the density of earth.

Statement-II: The time period of revolution of the satellite is $T = 2\pi\sqrt{\frac{R_e}{g}}$ (for satellite very close to the earth surface), where R_e is radius of earth and g is acceleration due to gravity.

In the light of the above statements, choose the correct answer from the options given below :

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- (1) Statement-I is true but Statement-II is false
- (2) Statement-I is false but Statement-II is true
- (3) Both Statement-I and Statement-II are true
- (4) Both Statement-I and Statement-II are false

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Q.9 Consider two boxes containing ideal gases A and B such that their temperatures, pressures and number densities are same. The molecular size of A is half of that of B and mass of molecule A is four times that of B. If the collision frequency in gas B is 32×10^{18} /s then collision frequency in gas A is _____ /s.

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- (1) 8×10^{18} (2) 2×10^{18}
- (3) 32×10^{18} (4) 4×10^{18}

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Q.10 In parallax method for the determination of focal length of a concave mirror, the object should always be placed:

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- (1) between the focus (F) and the centre of curvature (C) of the mirror ONLY
- (2) between the pole (P) and the focus (F) of the concave mirror ONLY
- (3) at any point beyond the focus (F) of the mirror
- (4) beyond the centre of the curvature (C) of the mirror ONLY

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Q.11 The wavelength of light, while it is passing through water is 540 nm. The refractive index of water is $4/3$. The wavelength of the same light when it is passing through a transparent medium having refractive index of $3/2$ is _____ nm.

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- (1) 540 (2) 840 (3) 480 (4) 380

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Q.12 The smallest wavelength of Lyman series is 91 nm. The difference between the largest wavelengths of Paschen and Balmer series is nearly _____ nm.

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- (1) 1784 (2) 1875 (3) 1217 (4) 1550

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- Q.13** An electric power line having total resistance of $2\ \Omega$, delivers 1 kW of power at 250 V . The percentage efficiency of transmission line is _____.

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- (1) 96.9 (2) 92.5 (3) 86.5 (4) 100

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- Q.14** Three small identical bubbles of water having same charge on each coalesce to form a bigger bubble. Then the ratio of the potentials on one initial bubble and that on the resultant bigger bubble is :

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- (1) $1:3^{2/3}$ (2) $3^{2/3}:1$ (3) $1:3^{1/3}$ (4) $1:2^{2/3}$

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- Q.15** Given below are two statements :

Statement-I: For a mechanical system of many particles total kinetic energy is the sum of kinetic energies of all the particles.

Statement-II: The total kinetic energy can be the sum of kinetic energy of the center of mass w.r.t. the origin and the kinetic energy of all the particles w.r.t. the center of mass as the reference.

In the light of the above statements, choose the correct answer from the options given below :

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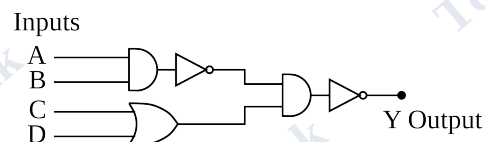
- (1) Both Statement-I and Statement-II are false
 (2) Statement-I is false but Statement-II is true
 (3) Statement-I is true but Statement-II is false
 (4) Both Statement-I and Statement-II are true

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- Q.16** The correct truth table for the given input data of the following logic gate is :

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Inputs				Output
A	B	C	D	Y
1	1	0	1	1
0	0	1	1	0
1	0	1	0	0
1	1	1	1	1

(1)

Inputs				Output
A	B	C	D	Y
1	1	0	1	0
0	0	1	1	1
1	0	1	0	1
1	1	1	1	1

(2)

Inputs				Output
A	B	C	D	Y
1	1	0	1	1
0	0	1	1	0
1	0	1	0	1
1	1	1	1	0

(3)

Inputs				Output
A	B	C	D	Y
1	1	0	1	0
0	0	1	1	0
1	0	1	0	1
1	1	1	1	1

(4)

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- Q.17** If ϵ , E and t represent the free space permittivity, electric field and time respectively, then the unit of $\frac{\epsilon E}{t}$ will be :

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- (1) Am (2) A/m (3) A/m^2 (4) Am^2

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- Q.18** Which of the following are true for a single slit diffraction?

- A. Width of central maxima increases with increase in wavelength keeping slit width constant.
 B. Width of central maxima increases with decrease in wavelength keeping slit width constant.
 C. Width of central maxima increases with decrease in slit width at constant wavelength.
 D. Width of central maxima increases with increase in slit width at constant wavelength.
 E. Brightness of central maxima increases for decrease in wavelength at constant slit width.

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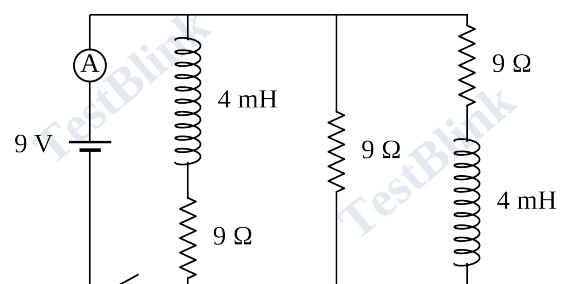
- (1) B, D only (2) A, D only
 (3) A, C, E only (4) B, C only

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- Q.19** Figure shows the circuit that contains three resistances ($9\ \Omega$ each) and two inductors (4 mH each). The reading of ammeter at the moment switch K is turned ON, is _____ A.

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- (1) 2 (2) 3 (3) 1 (4) zero

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Q.20 Light is incident on a metallic plate having work function 110×10^{-20} J. If the produced photoelectrons have zero kinetic energy then the angular frequency of the incident light is _____ rad/s. ($h = 6.63 \times 10^{-34}$ J s)

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- (1) 1.04×10^{13} (2) 1.66×10^{16}
(3) 1.66×10^{15} (4) 1.04×10^{16}

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Q.21 Among H_2S , H_2O , NF_3 , NH_3 and CHCl_3 , identify the molecule (X) with lowest dipole moment value. The number of lone pairs of electrons present on the central atom of the molecule (X) is :

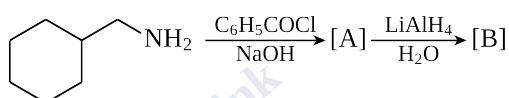
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- (1) 3 (2) 2 (3) 0 (4) 1

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Q.22



The final product [B] is :

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- (1)
- (2)
- (3)
- (4)

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Q.23 Given below are two statements :
Statement-I: $\text{C} < \text{O} < \text{N} < \text{F}$ is the correct order in terms of first ionization enthalpy values.
Statement-II: $\text{S} > \text{Se} > \text{Te} > \text{Po} > \text{O}$ is the correct order in terms of the magnitude of electron gain enthalpy values.

In the light of the above statements, choose the correct answer from the options given below :

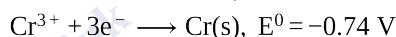
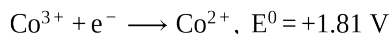
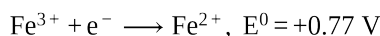
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- (1) Both Statement-I and Statement-II are false
(2) Statement-I is false but Statement-II is true
(3) Statement-I is true but Statement-II is false
(4) Both Statement-I and Statement-II are true

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Q.24 Consider the following reduction processes :



The tendency to act as reducing agent decreases in the order :

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- (1) $\text{Al} > \text{Fe}^{2+} > \text{Cr} > \text{Co}^{2+}$
(2) $\text{Al} > \text{Cr} > \text{Co}^{2+} > \text{Fe}^{2+}$
(3) $\text{Cr} > \text{Fe}^{2+} > \text{Al} > \text{Co}^{2+}$
(4) $\text{Al} > \text{Cr} > \text{Fe}^{2+} > \text{Co}^{2+}$

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Q.25 The compound A, $\text{C}_8\text{H}_8\text{O}_2$ reacts with acetophenone to form a single product via cross-Aldol condensation. The compound A on reaction with conc. NaOH forms a substituted benzyl alcohol as one of the two products. The compound A is :

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- (1) 2-hydroxy acetophenone
(2) 4-hydroxy benzaldehyde
(3) 4-methyl benzoic acid
(4) 4-methoxy benzaldehyde

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Q.26 Match List-I with List-II.

List-I Reaction of Glucose with		List-II Product formed	
A.	Hydroxylamine	I.	Gluconic acid
B.	Br_2 water	II.	Glucose pentaacetate
C.	Excess acetic anhydride	III.	Saccharic acid
D.	Concentrated HNO_3	IV.	Glucoxime

Choose the correct answer from the options given below :

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- (1) A-IV, B-III, C-II, D-I (2) A-I, B-III, C-IV, D-II
(3) A-III, B-I, C-IV, D-II (4) A-IV, B-I, C-II, D-III

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Q.27 Identify the correct statements :

Choose the correct answer from the options given below :

- A. Hydrated salts can be used as primary standard.
- B. Primary standard should not undergo any reaction with air.
- C. Reactions of primary standard with another substance should be instantaneous and stoichiometric.
- D. Primary standard should not be soluble in water.
- E. Primary standard should have low relative molar mass.

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- (1) D and E Only (2) A, B and E Only
(3) A, B, C and E Only (4) A, B and C Only

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Q.28 Correct statements regarding Arrhenius equation among the following are :

- A. Factor $e^{-E_a/RT}$ corresponds to fraction of molecules having kinetic energy less than E_a .
- B. At a given temperature, lower the E_a , faster is the reaction.
- C. Increase in temperature by about 10°C doubles the rate of reaction.
- D. Plot of $\log k$ vs $\frac{1}{T}$ gives a straight line with

$$\text{slope} = -\frac{E_a}{R}.$$

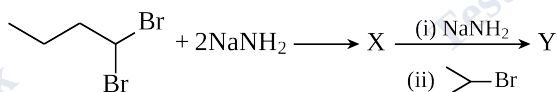
Choose the correct answer from the options given below :

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- (1) A and B Only (2) B and D Only
(3) B and C Only (4) A and C Only

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Q.29 Consider the following reaction :



The product Y formed is :

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- (1) 2-methylhex-3-yne (2) 5-methylhex-2-yne
(3) Isopropylbut-1-yne (4) 2-methylhex-2-yne

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Q.30 $[\text{Ni}(\text{PPh}_3)_2\text{Cl}_2]$ is a paramagnetic complex. Identify the INCORRECT statements about this complex.

- A. The complex exhibits geometrical isomerism.
- B. The complex is white in colour.
- C. The calculated spin-only magnetic moment of the complex is 2.84 BM.
- D. The calculated CFSE (Crystal Field Stabilization Energy) of Ni in this complex is $-0.8\Delta_o$.

E. The geometrical arrangement of ligands in this complex is similar to that in $\text{Ni}(\text{CO})_4$.

Choose the correct answer from the options given below :

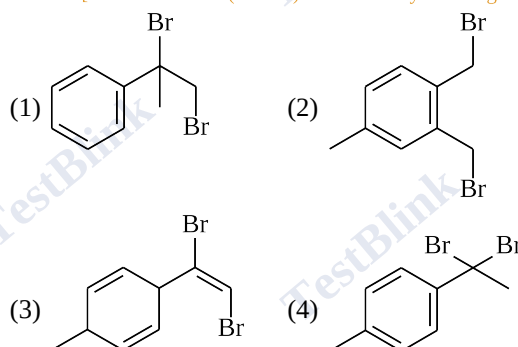
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- (1) C and D Only (2) A, B and D Only
(3) C, D and E Only (4) A and B Only

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Q.31 The dibromo compound [P] (molecular formula: $\text{C}_9\text{H}_{10}\text{Br}_2$) when heated with excess sodamide followed by treatment with dilute HCl gives [Q]. On warming [Q] with mercuric sulphate and dilute sulphuric acid yields [R] which gives positive Iodoform test but negative Tollens' test. The compound [P] is :

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Q.32 Which of the following mixture gives a buffer solution with $\text{pH} = 9.25$?

Given : $\text{pK}_b(\text{NH}_4\text{OH}) = 4.75$

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- (1) 0.5 M NH_4OH (0.2 L) + 0.2 M HCl (0.5 L)
(2) 0.2 M NH_4OH (0.5 L) + 0.1 M HCl (0.5 L)
(3) 0.2 M NH_4OH (0.4 L) + 0.1 M HCl (1 L)
(4) 0.4 M NH_4OH (1 L) + 0.1 M HCl (1 L)

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Q.33 When 1 g of compound (X) is subjected to Kjeldahl's method for estimation of nitrogen, 15 mL 1 M H_2SO_4 was neutralized by ammonia evolved. The percentage of nitrogen in compound (X) is :

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- (1) 0.21 (2) 21 (3) 42 (4) 0.42

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Q.34 The energy of first (lowest) Balmer line of H atom is x J. The energy (in J) of second Balmer line of H atom is :

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- (1) $\frac{x}{1.35}$ (2) $1.35x$ (3) x^2 (4) $2x$

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Q.35 Given below are two statements :

Statement-I: The first ionization enthalpy of Cr is lower than that of Mn.

Statement-II: The second and third ionization enthalpies of Cr are higher than those of Mn.

In the light of the above statements, choose the correct answer from the options given below :

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- (1) Statement-I is true but Statement-II is false
(2) Both Statement-I and Statement-II are true
(3) Both Statement-I and Statement-II are false
(4) Statement-I is false but Statement-II is true

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Q.36 At T(K), 100 g of 98% H_2SO_4 (w/w) aqueous solution is mixed with 100 g of 49% H_2SO_4 (w/w) aqueous solution. What is the mole fraction of H_2SO_4 in the resultant solution?

(Given : Atomic mass H = 1 u; S = 32 u; O = 16 u).
(Assume that temperature after mixing remains constant)

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- (1) 0.9 (2) 0.663 (3) 0.1 (4) 0.337

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Q.37 3,3-Dimethyl-2-butanol cannot be prepared by :

- A. $\xrightarrow[\text{H}_3\text{O}^+]{\text{MeMgBr-dry ether}}$
- B. $\xrightarrow{\text{H}_2\text{O}/\text{H}^+}$
- C. $\xrightarrow[\text{(ii) NaBH}_4/\text{MeOH}]{\text{(i) O}_3/\text{Zn-H}_2\text{O}}$
- D. $\xrightarrow{\text{LiAlH}_4/\text{H}_3\text{O}^+}$
- E. $\xrightarrow{\text{H}_2\text{O, Hg}^{2+}/\text{H}^+}$

Choose the correct answer from the options given below :

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(1) B and E Only

(2) B, C and E Only

(3) B and C Only

(4) B Only

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Q.38 $\text{A} + 2\text{B} \longrightarrow \text{AB}_2$

36.0 g of 'A' (Molar mass : 60 g mol^{-1}) and 56.0 g of 'B' (Molar mass : 80 g mol^{-1}) are allowed to react.

Which of the following statements are correct ?

- A. 'A' is the limiting reagent.
B. 77.0 g of AB_2 is formed.
C. Molar mass of AB_2 is 140 g mol^{-1} .
D. 15.0 g of A is left unreacted after the completion of reaction.

Choose the correct answer from the options given below :

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- (1) A and C Only (2) A and B Only
(3) C and D Only (4) B and D Only

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Q.39 Given below are two statements :

Statement-I: Elements 'X' and 'Y' are the most and least electronegative elements, respectively among N, As, Sb and P. The nature of the oxides X_2O_3 and Y_2O_3 is acidic and amphoteric, respectively.

Statement-II: BCl_3 is covalent in nature and gets hydrolysed in water. It produces $[\text{B}(\text{OH})_4]^-$ and $[\text{B}(\text{H}_2\text{O})_6]^{3+}$ in aqueous medium.

In the light of the above statements, choose the correct answer from the options given below :

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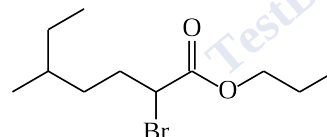
- (1) Statement-I is false but Statement-II is true
(2) Statement-I is true but Statement-II is false
(3) Both Statement-I and Statement-II are false
(4) Both Statement-I and Statement-II are true

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Q.40 The IUPAC name of the following compound is :

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- (1) 2-bromo-5-methylhexylpropanoate
(2) n-propyl-1-bromo-4-methylhexanoate
(3) n-propyl-2-bromo-5-methylheptanoate
(4) 2-bromo-5-methylpropanoate

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Q.41 Let $f(x) = [x]^2 - [x+3] - 3$, $x \in \mathbb{R}$, where $[.]$ is the greatest integer function. Then

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- (1) $f(x) = 0$ for finitely many values of x
 (2) $f(x) < 0$ only for $x \in [-1, 3]$
 (3) $\int_0^2 f(x) dx = -6$
 (4) $f(x) > 0$ only for $x \in [4, \infty)$

Scan or tap for the full solution



Q.42 Let the locus of the mid-point of the chord through the origin O of the parabola $y^2 = 4x$ be the curve S . Let P be any point on S . Then the locus of the point, which internally divides OP in the ratio $3 : 1$, is :

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- (1) $2x^2 = 3y$ (2) $2y^2 = 3x$ (3) $3y^2 = 2x$ (4) $3x^2 = 2y$

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Q.43 Let n be the number obtained on rolling a fair die. If the probability that the system $x - ny + z = 6$
 $x + (n - 2)y + (n + 1)z = 8$
 $(n - 1)y + z = 1$

has a unique solution is $\frac{k}{6}$, then the sum of k and all possible values of n is :

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- (1) 22 (2) 20 (3) 24 (4) 21

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Q.44 Let the domain of the function

$$f(x) = \log_3 \log_5 (7 - \log_2 (x^2 - 10x + 85)) + \sin^{-1} \left(\left| \frac{3x - 7}{17 - x} \right| \right)$$

be $(\alpha, \beta]$. Then $\alpha + \beta$ is equal to :

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- (1) 12 (2) 8 (3) 10 (4) 9

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Q.45 Let S and S' be the foci of the ellipse $\frac{x^2}{25} + \frac{y^2}{9} = 1$ and $P(\alpha, \beta)$ be a point on the ellipse in the first quadrant. If $(SP)^2 + (S'P)^2 - SP \cdot S'P = 37$, then $\alpha^2 + \beta^2$ is equal to :

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- (1) 13 (2) 15 (3) 11 (4) 17

Scan or tap for the full solution



Q.46 If $\lim_{x \rightarrow 0} \frac{e^{(a-1)x} + 2\cos bx + (c-2)e^{-x}}{x\cos x - \log_e(1+x)} = 2$, then $a^2 + b^2 + c^2$ is equal to :

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- (1) 3 (2) 5 (3) 9 (4) 7

Scan or tap for the full solution



Q.47 Let $\vec{a} = 2\hat{i} - \hat{j} + \hat{k}$ and $\vec{b} = \lambda\hat{j} + 2\hat{k}$, $\lambda \in \mathbb{Z}$ be two vectors. Let $\vec{c} = \vec{a} \times \vec{b}$ and $\vec{d}$ be a vector of magnitude 2 in yz -plane. If $|\vec{c}| = \sqrt{53}$, then the maximum possible value of $(\vec{c} \cdot \vec{d})^2$ is equal to :

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- (1) 52 (2) 26 (3) 104 (4) 208

Scan or tap for the full solution



Q.48 Let $[\cdot]$ denote the greatest integer function, and let $f(x) = \min \{ \sqrt{2}x, x^2 \}$. Let $S = \{x \in (-2, 2) : \text{the function } g(x) = |x|[x^2] \text{ is discontinuous at } x\}$.

Then $\sum_{x \in S} f(x)$ equals

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- (1) $2 - \sqrt{2}$ (2) $2\sqrt{6} - 3\sqrt{2}$
 (3) $1 - \sqrt{2}$ (4) $\sqrt{6} - 2\sqrt{2}$

Scan or tap for the full solution



Q.49 Among the statements

(S1) : If $A(5, -1)$ and $B(-2, 3)$ are two vertices of a triangle, whose orthocentre is $(0, 0)$, then its third vertex is $(-4, -7)$

and

(S2) : If positive numbers $2a, b, c$ are three consecutive terms of an A.P., then the lines $ax + by + c = 0$ are concurrent at $(2, -2)$,

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- (1) both are incorrect (2) only (S2) is correct
 (3) both are correct (4) only (S1) is correct

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Q.50 Let α, β be the roots of the quadratic equation $12x^2 - 20x + 3\lambda = 0$, $\lambda \in \mathbb{Z}$. If $\frac{1}{2} \leq |\beta - \alpha| \leq \frac{3}{2}$, then the sum of all possible values of λ is :

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- (1) 3 (2) 6 (3) 1 (4) 4

Scan or tap for the full solution



Q.51 The number of elements in the relation $\text{mathrmR} = \{(x, y) : 4x^2 + y^2 < \text{mathrm52}, x, y \text{ in } \text{mathbfZ}\}$ is

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- (1) 86 (2) 67 (3) 89 (4) 77

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- Q.52** Let f and g be functions satisfying $f(x+y) = f(x)f(y)$, $f(1) = 7$ and $g(x+y) = g(xy)$, $g(1) = 1$, for all $x, y \in \mathbb{N}$. If $\sum_{x=1}^n \left(\frac{f(x)}{g(x)} \right) = 19607$, then n is equal to : [JEE Main 2026 (Online) 22nd January Evening Shift]
- (1) 6 (2) 7 (3) 4 (4) 5

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- Q.53** If $X = \begin{bmatrix} x \\ y \\ z \end{bmatrix}$ is a solution of the system of equations

$$AX = B, \text{ where } \text{adj } A = \begin{bmatrix} 4 & 2 & 2 \\ -5 & 0 & 5 \\ 1 & -2 & 3 \end{bmatrix} \text{ and } B = \begin{bmatrix} 4 \\ 0 \\ 2 \end{bmatrix},$$

then $|x+y+z|$ is equal to :

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- (1) 3 (2) 2 (3) $\frac{3}{2}$ (4) 1

Scan or tap for the full solution



- Q.54** The area of the region

$$A = \{(x, y) : 4x^2 + y^2 \leq 8 \text{ and } y^2 \leq 4x\} \text{ is:}$$

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- (1) $\pi + \frac{2}{3}$ (2) $\frac{\pi}{2} + 2$ (3) $\pi + 4$ (4) $\frac{\pi}{2} + \frac{1}{3}$

Scan or tap for the full solution



- Q.55** Let L be the line $\frac{x+1}{2} = \frac{y+1}{3} = \frac{z+3}{6}$ and let S be the set of all points (a, b, c) on L , whose distance from the line $\frac{x+1}{2} = \frac{y+1}{3} = \frac{z-9}{0}$ along the line L is 7. Then $\sum_{(a, b, c) \in S} (a+b+c)$ is equal to :

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- (1) 28 (2) 6 (3) 40 (4) 34

Scan or tap for the full solution



- Q.56** If $y = y(x)$ satisfies the differential equation

$$16(\sqrt{9x+x\sqrt{x}})(4+\sqrt{9+\sqrt{x}})\cos y \, dy = (1+2\sin y) \, dx, \quad x > 0 \text{ and } y(256) = \frac{\pi}{2}, \quad y(49) = \alpha, \text{ then } 2\sin \alpha \text{ is equal to}$$

: [JEE Main 2026 (Online) 22nd January Evening Shift]

- (1) $2\sqrt{2} - 1$ (2) $\sqrt{2} - 1$
(3) $2(\sqrt{2} - 1)$ (4) $3(\sqrt{2} - 1)$

Scan or tap for the full solution



- Q.57** Let $P(10, 2\sqrt{15})$ be a point on the hyperbola $\frac{x^2}{a^2} - \frac{y^2}{b^2} = 1$, whose foci are S and S' . If the length of its latus rectum is 8, then the square of the area of $\Delta PSS'$ is equal to :

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- (1) 4200 (2) 1462 (3) 900 (4) 2700

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- Q.58** If the mean deviation about the median of the numbers $k, 2k, 3k, \dots, 1000k$ is 500, then k^2 is equal to :

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- (1) 1 (2) 16 (3) 9 (4) 4

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- Q.59** Let $S = \{z \in \mathbb{C} : 4z^2 + \bar{z} = 0\}$. Then $\sum_{z \in S} |z|^2$ is equal to :

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- (1) $\frac{5}{64}$ (2) $\frac{1}{16}$ (3) $\frac{7}{64}$ (4) $\frac{3}{16}$

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- Q.60** Let C_r denote the coefficient of x^r in the binomial expansion of $(1+x)^n$, $n \in \mathbb{N}$, $0 \leq r \leq n$. If

$$P_n = C_0 - C_1 + \frac{2^2}{3}C_2 - \frac{2^3}{4}C_3 + \dots + \frac{(-2)^n}{n+1}C_n, \text{ then the value of } \sum_{n=1}^{25} \frac{1}{P_{2n}} \text{ equals}$$

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- (1) 675 (2) 580 (3) 525 (4) 650

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Section B (Numerical) — 15 questions

- Q.61** An insulated cylinder of volume 60 cm^3 is filled with a gas at 27°C and 2 atmospheric pressure. Then the gas is compressed making the final volume as 20 cm^3 while allowing the temperature to rise to 77°C . The final pressure is _____ atmospheric pressure.

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- Q.62** The mass of benzanilide obtained from the benzoylation reaction of 5.8 g of aniline, if yield of product is 82%, is _____ g (nearest integer). (Given molar mass in g mol^{-1} H : 1, C : 12, N : 14, O : 16)

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- Q.63** A conducting circular loop is rotated about its diameter at a constant angular speed of 100 rad/s in a magnetic field of 0.5 T perpendicular to the axis of rotation. When the loop is rotated by 30° from the horizontal position, the induced EMF is 15.4 mV. The radius of the loop is _____ mm.

(Take $\pi = \frac{22}{7}$)

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Scan or tap for the full solution



- Q.64** A capacitor P with capacitance 10×10^{-6} F is fully charged with a potential difference of 6.0 V and disconnected from the battery. The charged capacitor P is connected across another capacitor Q with capacitance 20×10^{-6} F. The charge on capacitor Q when equilibrium is established will be $\alpha \times 10^{-5}$ C (assume capacitor Q does not have any charge initially), the value of α is _____.

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Scan or tap for the full solution



- Q.65** A cylindrical conductor of length 2 m and area of cross-section 0.2 mm^2 carries an electric current of 1.6 A when its ends are connected to a 2 V battery. Mobility of electrons in the conductor is $\alpha \times 10^{-3} \text{ m}^2/\text{V s}$. The value of α is : (electron concentration = $5 \times 10^{28} / \text{m}^3$ and electron charge = $1.6 \times 10^{-19} \text{ C}$)

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- Q.66** Two masses m and 2m are connected by a light string going over a pulley (disc) of mass 30m with radius $r = 0.1 \text{ m}$. The pulley is mounted in a vertical plane and it is free to rotate about its axis. The 2m mass is released from rest and its speed when it has descended through a height of 3.6 m is _____ m/s. (Assume string does not slip and $g = 10 \text{ m/s}^2$)

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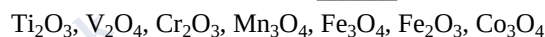
- Q.67** Consider the following electrochemical cell :
 $\text{Pt} | \text{O}_2(\text{g})(1 \text{ bar}) | \text{HCl}(\text{aq}) || \text{M}^{2+}(\text{aq}, 1.0 \text{ M}) | \text{M}(\text{s})$
 The pH above which, oxygen gas would start to evolve at anode is _____ (nearest integer).
 [Given : $E^\circ_{\text{M}^{2+}/\text{M}} = 0.994 \text{ V}$
 $E^\circ_{\text{O}_2/\text{H}_2\text{O}} = 1.23 \text{ V}$ } standard reduction potential and $\frac{RT}{F}(2.303) = 0.059 \text{ V}$ at the given condition]

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- Q.68** Among the following oxides of 3d elements, the number of mixed oxides are _____ .



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- Q.69** If the enthalpy of sublimation of Li is 155 kJ mol^{-1} , enthalpy of dissociation of F_2 is 150 kJ mol^{-1} , ionization enthalpy of Li is 520 kJ mol^{-1} , electron gain enthalpy of F is -313 kJ mol^{-1} , standard enthalpy of formation of LiF is -594 kJ mol^{-1} . The magnitude of lattice enthalpy of LiF is _____ kJ mol^{-1} . (Nearest Integer)

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- Q.70** Consider $\text{A} \xrightarrow{k_1} \text{B}$ and $\text{C} \xrightarrow{k_2} \text{D}$ are two reactions. If the rate constant (k_1) of the $\text{A} \rightarrow \text{B}$ reaction can be expressed by the following equation

$$\log_{10} k = 14.34 - \frac{1.5 \times 10^4}{T/K} \text{ and activation energy of}$$

$\text{C} \rightarrow \text{D}$ reaction (E_{a_2}) is $\frac{1}{5}$ th of the $\text{A} \rightarrow \text{B}$ reaction (E_{a_1}), then the value of (E_{a_2}) is _____ kJ mol^{-1} . (Nearest Integer)

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- Q.71** Suppose a, b, c are in A.P. and $a^2, 2b^2, c^2$ are in G.P. If $a < b < c$ and $a + b + c = 1$, then $9(a^2 + b^2 + c^2)$ is equal to _____.

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- Q.72** Let a vector $\vec{a} = \sqrt{2}\hat{i} - \hat{j} + \lambda\hat{k}$, $\lambda > 0$, make an obtuse angle with the vector $\vec{b} = -\lambda^2\hat{i} + 4\sqrt{2}\hat{j} + 4\sqrt{2}\hat{k}$ and an angle θ , $\frac{\pi}{6} < \theta < \frac{\pi}{2}$, with the positive z-axis. If the set of all possible values of λ is $(\alpha, \beta) - \{\gamma\}$, then $\alpha + \beta + \gamma$ is equal to _____.

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- Q.73** Let $[.]$ be the greatest integer function. If $\alpha = \int_0^{64} (x^{1/3} - [x^{1/3}]) dx$, then $\frac{1}{\pi} \int_0^{\alpha\pi} \left(\frac{\sin^2 \theta}{\sin^6 \theta + \cos^6 \theta} \right) d\theta$ is equal to _____.

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- Q.74** Let $\cos(\alpha + \beta) = -\frac{1}{10}$ and $\sin(\alpha - \beta) = \frac{3}{8}$, where $0 < \alpha < \frac{\pi}{3}$ and $0 < \beta < \frac{\pi}{4}$. If $\tan 2\alpha = \frac{3(1-r\sqrt{5})}{\sqrt{11(s+\sqrt{5})}}$, $r, s \in \mathbb{N}$, then $r + s$ is equal to _____.

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- Q.75** Let S be the set of the first 11 natural numbers. Then the number of elements in $A = \{B \subseteq S : n(B) \geq 2 \text{ and the product of all elements of } B \text{ is even}\}$ is _____.

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ANSWER KEY

Section A (MCQ) — 60 questions

Ques.	1	2	3	4	5	6	7	8	9	10	11
Ans.	2	1	1	1	2	3	2	3	4	3	3
Ques.	12	13	14	15	16	17	18	19	20	21	22
Ans.	3	1	1	4	1	3	3	3	4	4	1
Ques.	23	24	25	26	27	28	29	30	31	32	33
Ans.	4	4	4	4	4	3	1	2	4	2	3
Ques.	34	35	36	37	38	39	40	41	42	43	44
Ans.	2	1	4	1	4	2	3	2	2	1	4
Ques.	45	46	47	48	49	50	51	52	53	54	55
Ans.	1	4	4	3	3	1	4	4	2	1	4
Ques.	56	57	58	59	60						
Ans.	1	4	4	4	1						

Section B (Numerical) — 15 questions

Ques.	61	62	63	64	65	66	67
Ans.	7	10	14	4	1	2	4
Ques.	68	69	70	71	72	73	74
Ans.	3	1031	57	9	5	36	20
Ques.	75						
Ans.	1979						