

NEET AMINES — BASICITY OF AMINES — PYQS

Max. Marks: 28

Questions: 7

Section A (MCQ) — 7 questions

- Q.1** Given below are two statements:
In light of the above statements, choose the most appropriate answer from the options given below.

Statement-I: Oxidation of p-nitrotoluene with acidic KMnO_4 gives an acid that is stronger than benzoic acid.

Statement-II: Reduction of p-nitrotoluene with Sn/HCl followed by neutralization gives an amine that is more basic than aniline.

[RE-NEET (UG) 2026]

- (1) Statement-I is incorrect but Statement-II is correct
- (2) Both Statement-I and Statement-II are correct
- (3) Both Statement-I and Statement-II are incorrect
- (4) Statement-I is correct but Statement-II is incorrect

 Scan or tap for the
full solution


- Q.2** The correct order of decreasing basic strength of the given amines is:

[NEET (UG) 2025]

- (1) N-ethylethanamine > ethanamine > N-methylaniline > benzenamine
- (2) benzenamine > ethanamine > N-methylaniline > N-ethylethanamine
- (3) N-methylaniline > benzenamine > ethanamine > N-ethylethanamine
- (4) N-ethylethanamine > ethanamine > benzenamine > N-methylaniline

 Scan or tap for the
full solution


- Q.3** The correct order of the basic strength of methyl substituted amines in aqueous solution is :

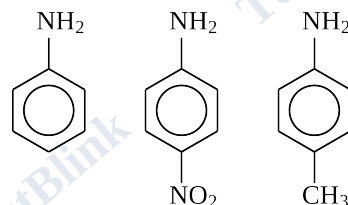
[NEET (UG) 2019]

- (1) $(\text{CH}_3)_3\text{N} > (\text{CH}_3)_2\text{NH} > \text{CH}_3\text{NH}_2$
- (2) $\text{CH}_3\text{NH}_2 > (\text{CH}_3)_2\text{NH} > (\text{CH}_3)_3\text{N}$
- (3) $(\text{CH}_3)_2\text{NH} > \text{CH}_3\text{NH}_2 > (\text{CH}_3)_3\text{N}$
- (4) $(\text{CH}_3)_3\text{N} > \text{CH}_3\text{NH}_2 > (\text{CH}_3)_2\text{NH}$

 Scan or tap for the
full solution


- Q.4** The correct increasing order of basic strength for the following compounds is

[NEET 2017]



(I)

(II)

(III)

(1) $\text{III} < \text{I} < \text{II}$ (2) $\text{III} < \text{II} < \text{I}$ (3) $\text{II} < \text{I} < \text{III}$ (4) $\text{II} < \text{III} < \text{I}$
 Scan or tap for the
full solution


- Q.5** The correct statement regarding the basicity of arylamines is

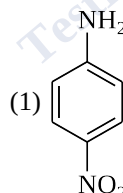
[NEET 2016 (Phase 1)]

- (1) arylamines are generally more basic than alkylamines because of aryl group
- (2) arylamines are generally more basic than alkylamines, because the nitrogen atom in arylamines is sp -hybridised
- (3) arylamines are generally less basic than alkylamines because the nitrogen lone-pair electrons are delocalised by interaction with the aromatic ring π -electron system
- (4) arylamines are generally more basic than alkylamines because the nitrogen lone-pair electrons are not delocalised by interaction with the aromatic ring π -electron system.

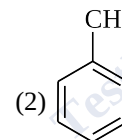
 Scan or tap for the
full solution


- Q.6** Which of the following compounds is most basic?

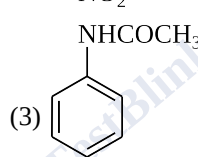
[AIPMT 2011 (Mains)]



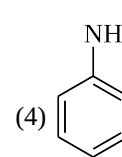
(1)



(2)



(3)



(4)

 Scan or tap for the
full solution


Q.7 Which of the following is more basic than aniline?

[AIPMT 2006]

- (1) Benzylamine (2) Diphenylamine
(3) Triphenylamine (4) p-Nitroaniline

Scan or tap for the
full solution



ANSWER KEY

Ques.	1	2	3	4	5	6	7				
Ans.	2	1	3	3	3	2	1				