

NEET COORDINATION COMPOUNDS — NOMENCLATURE AND FORMULA WRITING OF COORDINATION COMPOUNDS — PYQS

Max. Marks: 24

Questions: 6

Section A (MCQ) — 6 questions

Q.1 The formula of tetraammineaquachloridocobalt(III) chloride is [RE-NEET (UG) 2026]

- (1) $[\text{Co}(\text{NH}_3)_4(\text{H}_2\text{O})\text{Cl}]\text{Cl}_2$
- (2) $[\text{Co}(\text{NH}_3)_4\text{Cl}_2] \cdot \text{H}_2\text{O}$
- (3) $[\text{Co}(\text{NH}_3)_4]\text{Cl}_3 \cdot \text{H}_2\text{O}$
- (4) $[\text{Co}(\text{NH}_3)_4(\text{H}_2\text{O})\text{Cl}]\text{Cl}$

Scan or tap for the full solution



Q.2 The IUPAC name of the complex $[\text{Ag}(\text{H}_2\text{O})_2][\text{Ag}(\text{CN})_2]$ is: [NEET (UG) 2022]

- (1) dicyanidosilver(II) diaquaargentate(II)
- (2) diaquasilver(II) dicyanidoargentate(II)
- (3) dicyanidosilver(I) diaquaargentate(I)
- (4) diaquasilver(I) dicyanidoargentate(I)

Scan or tap for the full solution



Q.3 The name of complex ion, $[\text{Fe}(\text{CN})_6]^{3-}$ is [AIPMT 2015]

- (1) hexacyanoferrate (III) ion
- (2) tricyanoferrate (III) ion
- (3) hexacyanidoferrate (III) ion
- (4) hexacyanoiron (III) ion.

Scan or tap for the full solution



Q.4 The correct IUPAC name for $[\text{CrF}_2(\text{en})_2]\text{Cl}$ is [NEET 2013 (Karnataka)]

- (1) chloro difluorido ethylene diaminechromium (III) chloride
- (2) difluoridobis (ethylene diamine) chromium (III) chloride
- (3) difluorobis-(ethylene diamine) chromium (III) chloride
- (4) chloro difluoridobis (ethylene diamine) chromium (III)

Scan or tap for the full solution



Q.5 The hypothetical complex chloro diaquatriammine cobalt(III) chloride can be represented as [AIPMT 2002]

- (1) $[\text{CoCl}(\text{NH}_3)_3(\text{H}_2\text{O})_2]\text{Cl}_2$
- (2) $[\text{Co}(\text{NH}_3)_3(\text{H}_2\text{O})\text{Cl}_3]$
- (3) $[\text{Co}(\text{NH}_2)_3(\text{H}_2\text{O})_2\text{Cl}]$
- (4) $[\text{Co}(\text{NH}_3)_3(\text{H}_2\text{O})_3]\text{Cl}_3$

Scan or tap for the full solution



Q.6 CuSO_4 when reacts with KCN forms CuCN , which is insoluble in water. It is soluble in excess of KCN, due to formation of the following complex [AIPMT 2002]

- (1) $\text{K}_2[\text{Cu}(\text{CN})_4]$
- (2) $\text{K}_3[\text{Cu}(\text{CN})_4]$
- (3) CuCN_2
- (4) $\text{Cu}[\text{KC}(\text{CN})_4]$

Scan or tap for the full solution



ANSWER KEY

Ques.	1	2	3	4	5	6					
Ans.	1	4	3	2	1	2					