

## NEET COORDINATION COMPOUNDS — WERNER'S THEORY — PYQS

Max. Marks: 24

Questions: 6

## Section A (MCQ) — 6 questions

**Q.1** A 1:3 electrolyte in an aqueous solution is  
 [RE-NEET (UG) 2026]

- (1)  $[\text{Co}(\text{NH}_3)_3(\text{NO}_2)_3]$  (2)  $[\text{CoCl}_2(\text{NH}_3)_4]\text{Cl}$   
 (3)  $[\text{CoCl}(\text{NH}_3)_5]\text{Cl}_2$  (4)  $[\text{Co}(\text{NH}_3)_6]\text{Cl}_3$

 Scan or tap for the  
 full solution


**Q.2** Out of the following complex compounds, which of the compound will be having the minimum conductance solution?  
 [NEET (UG) 2025]

- (1)  $[\text{Co}(\text{NH}_3)_6]\text{Cl}_3$  (2)  $[\text{Co}(\text{NH}_3)_5\text{Cl}]\text{Cl}$   
 (3)  $[\text{Co}(\text{NH}_3)_3\text{Cl}_3]$  (4)  $[\text{Co}(\text{NH}_3)_4\text{Cl}_2]$

 Scan or tap for the  
 full solution


**Q.3** The correct order of the stoichiometries of  $\text{AgCl}$  formed when  $\text{AgNO}_3$  in excess is treated with the complexes :  $\text{CoCl}_3 \cdot 6\text{NH}_3$ ,  $\text{CoCl}_3 \cdot 5\text{NH}_3$ ,  $\text{CoCl}_3 \cdot 4\text{NH}_3$  respectively is  
 [NEET 2017]

- (1) 3AgCl, 1AgCl, 2AgCl  
 (2) 3AgCl, 2AgCl, 1AgCl  
 (3) 2AgCl, 3AgCl, 2AgCl  
 (4) 1AgCl, 3AgCl, 2AgCl

 Scan or tap for the  
 full solution


**Q.4** Cobalt (III) chloride forms several octahedral complexes with ammonia. Which of the following will not give test for chloride ions with silver nitrate at  $25^\circ\text{C}$ ?  
 [AIPMT 2015 (Cancelled)]

- (1)  $\text{CoCl}_3 \cdot 5\text{NH}_3$  (2)  $\text{CoCl}_3 \cdot 6\text{NH}_3$   
 (3)  $\text{CoCl}_3 \cdot 3\text{NH}_3$  (4)  $\text{CoCl}_3 \cdot 4\text{NH}_3$

 Scan or tap for the  
 full solution


**Q.5** An excess of  $\text{AgNO}_3$  is added to 100 mL of a 0.01 M solution of dichlorotetraaquachromium(III) chloride. The number of moles of  $\text{AgCl}$  precipitated would be  
 [NEET 2013]

- (1) 0.003 (2) 0.01 (3) 0.001 (4) 0.002

 Scan or tap for the  
 full solution


**Q.6** Which of the following will exhibit maximum ionic conductivity?  
 [AIPMT 2001]

- (1)  $\text{K}_4[\text{Fe}(\text{CN})_6]$  (2)  $[\text{Co}(\text{NH}_3)_6]\text{Cl}_3$   
 (3)  $[\text{Cu}(\text{NH}_3)_4]\text{Cl}_2$  (4)  $[\text{Ni}(\text{CO})_4]$

 Scan or tap for the  
 full solution


## ANSWER KEY

| Ques. | 1 | 2    | 3 | 4 | 5 | 6 |  |
|-------|---|------|---|---|---|---|--|
| Ans.  | 4 | 3, 4 | 2 | 3 | 3 | 1 |  |