

NEET D- AND F-BLOCK ELEMENTS — GENERAL CHARACTERISTICS, CHEMICAL PROPERTIES AND USES OF LANTHANOIDS — PYQS

Max. Marks: 12

Questions: 3

Section A (MCQ) — 3 questions

Q.1 The UV-visible absorption bands in the spectra of lanthanoid ions are 'X', probably because of the excitation of electrons involving 'Y'. The 'X' and 'Y', respectively, are : [RE-NEET (UG) 2024]

- (1) Broad and f orbitals
- (2) Narrow and f orbitals
- (3) Broad and d and f orbitals
- (4) Narrow and d and f orbitals

Scan or tap for the full solution



Q.2 Gadolinium has a low value of third ionisation enthalpy because of [NEET (UG) 2022]

- (1) small size
- (2) high exchange enthalpy
- (3) high electronegativity
- (4) high basic character

Scan or tap for the full solution



Q.3 Which one of the following statements related to lanthanoids is incorrect? [NEET 2016 (Phase 2)]

- (1) Europium shows +2 oxidation state.
- (2) The basicity decreases as the ionic radius decreases from Pr to Lu.
- (3) All the lanthanons are much more reactive than aluminium.
- (4) Ce(+4) solutions are widely used as oxidizing agent in volumetric analysis.

Scan or tap for the full solution



ANSWER KEY

Ques.	1	2	3									
Ans.	2	2	3									