

NEET D- AND F-BLOCK ELEMENTS — POSITION AND ELECTRONIC CONFIGURATION OF D-BLOCK ELEMENTS — PYQS

Max. Marks: 12

Questions: 3

Section A (MCQ) — 3 questions

Q.1 Sc ($Z = 21$) is a transition element but Zn ($Z = 30$) is not because [NEET 2013 (Karnataka)]

- (1) both Sc^{3+} and Zn^{2+} ions are colourless and form white compounds.
- (2) in case of Sc, 3d orbitals are partially filled but in Zn these are filled.
- (3) last electron is assumed to be added to 4s level in case of Zn.
- (4) both Sc and Zn do not exhibit variable oxidation states.

Scan or tap for the full solution



Q.2 Which of the following ions has electronic configuration $[\text{Ar}]3d^6$? (At. nos. Mn = 25, Fe = 26, Co = 27, Ni = 28)

[AIPMT 2010 (Prelims)]

- (1) Ni^{3+} (2) Mn^{3+} (3) Fe^{3+} (4) Co^{3+}

Scan or tap for the full solution



Q.3 Among the following series of transition metal ions, the one where all metal ions have $3d^2$ electronic configuration is

[At. Nos. Ti = 22, V = 23, Cr = 24, Mn = 25]

[AIPMT 2004]

- (1) Ti^{3+} , V^{2+} , Cr^{3+} , Mn^{4+} (2) Ti^{+} , V^{4+} , Cr^{6+} , Mn^{7+}
 (3) Ti^{4+} , V^{3+} , Cr^{2+} , Mn^{3+} (4) Ti^{2+} , V^{3+} , Cr^{4+} , Mn^{5+}

Scan or tap for the full solution



ANSWER KEY

Ques.	1	2	3								
Ans.	2	4	4								