

NEET SOLUTIONS — TYPES OF SOLUTIONS AND CONCENTRATION TERMS — PYQS

Max. Marks: 20
Questions: 5

Section A (MCQ) — 5 questions

Q.1 Identify the correct statements :
Choose the correct answer from the options given below :

- A. The molality of 2.5 g of ethanoic acid (Molar mass : 60 g mol^{-1}) in 75 g of benzene solution is 0.556 m.
- B. The molarity of a solution containing 5 g of NaOH (molar mass : 40 g mol^{-1}) in 450 mL of solution is 0.278 M at 298 K.
- C. Aquatic species are more comfortable in cold water.
- D. The solubility of gas increases with decrease in pressure.
- E. For a binary mixture of A and B, the number of moles of A and B are n_A and n_B respectively.

The mole fraction of B will be $x_B = \frac{n_B}{n_A + n_B}$.

[NEET (UG) 2026]

- (1) A, B and C only (2) A, D and E only
(3) A and B only (4) A and C only

Scan or tap for the full solution



Q.2 Match List-I with List-II

List - I (Example)		List - II (Type of Solution)	
A.	Humidity	I.	Solid in solid
B.	Alloys	II.	Liquid in gas
C.	Amalgams	III.	Solid in gas
D.	Smoke	IV.	Liquid in solid

Choose the correct answer from the options given below:

[NEET (UG) 2025]

- (1) A-III, B-I, C-IV, D-II (2) A-III, B-II, C-I, D-IV
(3) A-II, B-IV, C-I, D-III (4) A-II, B-I, C-IV, D-III

Scan or tap for the full solution



Q.3 What is the mole fraction of the solute in a 1.00 m aqueous solution?

[AIPMT 2015]

- (1) 1.770 (2) 0.0354 (3) 0.0177 (4) 0.177

Scan or tap for the full solution



Q.4 Which of the following is dependent on temperature?
[NEET 2017]

- (1) Molarity (2) Mole fraction
(3) Weight percentage (4) Molality

Scan or tap for the full solution



Q.5 The mole fraction of the solute in one molal aqueous solution is
[AIPMT 2005]

- (1) 0.009 (2) 0.018 (3) 0.027 (4) 0.036

Scan or tap for the full solution



ANSWER KEY

Ques.	1	2	3	4	5						
Ans.	1	4	3	1	2						