

NEET STRUCTURE OF ATOM — SHAPES OF ATOMIC ORBITALS — PYQS

Max. Marks: 8

Questions: 2

Section A (MCQ) — 2 questions

Q.1 Identify the incorrect statement from the following.

[NEET (UG) 2022]

- (1) All the five 5d orbitals are different in size when compared to the respective 4d orbitals.
- (2) All the five 4d orbitals have shapes similar to the respective 3d orbitals.
- (3) In an atom, all the five 3d orbitals are equal in energy in free state.
- (4) The shapes of d_{xy} , d_{yz} and d_{zx} orbitals are similar to each other; and $d_{x^2-y^2}$ and d_{z^2} are similar to each other.

Scan or tap for the full solution



Q.2 Which of the following pairs of d-orbitals will have electron density along the axes ?

[NEET 2016 (Phase 2)]

- | | |
|-------------------------------|------------------------------|
| (1) d_{z^2} , d_{xz} | (2) d_{xz} , d_{yz} |
| (3) d_{z^2} , $d_{x^2-y^2}$ | (4) d_{xy} , $d_{x^2-y^2}$ |

Scan or tap for the full solution



ANSWER KEY

Ques.	1	2									
Ans.	4	3									