

NEET THERMODYNAMICS (PHYSICS) — SPECIFIC HEAT CAPACITIES OF GASES (CP, CV) — PYQS

Max. Marks: 12

Questions: 3

Section A (MCQ) — 3 questions

- Q.1** The molar specific heats of an ideal gas at constant pressure and volume are denoted by C_p and C_v , respectively. If $\gamma = \frac{C_p}{C_v}$ and R is the universal gas constant, then C_v is equal to [NEET 2013]

(1) $\frac{(\gamma - 1)}{R}$ (2) γR (3) $\frac{1 + \gamma}{1 - \gamma}$ (4) $\frac{R}{(\gamma - 1)}$

Scan or tap for the full solution



- Q.2** If c_p and c_v denote the specific heats (per unit mass) of an ideal gas of molecular weight M , then [AIPMT 2010 (Mains)]

(1) $c_p - c_v = \frac{R}{M^2}$ (2) $c_p - c_v = R$
 (3) $c_p - c_v = \frac{R}{M}$ (4) $c_p - c_v = MR$

Scan or tap for the full solution



- Q.3** The molar specific heat at constant pressure of an ideal gas is $(7/2) R$. The ratio of specific heat at constant pressure to that at constant volume is

[AIPMT 2006]

(1) $9/7$ (2) $7/5$ (3) $8/7$ (4) $5/7$

Scan or tap for the full solution



ANSWER KEY

Ques.	1	2	3								
Ans.	4	3	2								